GENERAL CHEMISTRY
CHEM 1010-03
AUTUMN 2010

Instructor: Dr. Martin Margittai
SGM 253
Phone: 303-871-4135
Email: martin.margittai@du.edu

Text: Chemistry, 5th Edition, Silberberg

Lectures: 8-8:50 am, MWF, Olin 205
Discussion: 8-8:50 am, R, Olin 205
Office Hours: 2-3:00 pm, MF, SGM253

Exams: There are 3 1-hour midterm exams during the quarter, plus a 2-hour cumulative final exam. Each exam is worth 100 points. Exam questions will be similar to the problems worked in class.

If you miss a 1-hour midterm exam, then your final exam will be counted twice and replace the missed midterm exam. With one exception, **THERE WILL BE NO MAKEUP EXAMS**. The only exception to the no-makeup policy will be for members of a university team or group, e.g. athletic team or music group scheduled to be away from campus at the time of the exam. You must inform your instructor of this prior to the exam and make arrangements at that time for a makup exam.

If you take all 3 midterm exams and your grade on the final exam is better than one of your midterm exam grades, then your final exam will be counted twice and replace your lowest midterm exam grade.

There are 10 quizzes that will be administered during discussion periods. Your top 8 scores will count towards your grade, i.e. your 2 lowest quiz grades will be dropped. Clicker questions will be given during lecture starting from week 2. Points for those questions will be given based on participation (50%) and correctness (50%).

Grading:		
	midterm exams	300 points
	final exam	100 points
	quizzes	120 points
	clicker questions	80 points

The assignment of a letter grade to a given numerical grade will depend on the overall class performance. However, if everybody does well, grades will not be curved down. **Also, note that points will be deducted from your final grade for disruptive behavior**.

DATE		TOPIC	READING
WEEK 1			
Sep	13	Intro. to class	Ch1-3
	15	Nature of light	Ch7.1-7.2
	16	Discussion, quiz	
	17	Wave-Particle/Quantum-Mechanical	Ch7.3-7.4
WEEK 2			
Sep	20	Periodic table	Ch8.1-8.2
	22	Atomic properties	Ch8.3-8.4
	23	Discussion, quiz	
	24	Chemical reactivity	Ch8.5
WEEK 3			
Sep	27	Chemical Bonds	Ch9.1-9.2
	29	Covalent bond	Ch9.3-9.6
	30	Discussion, quiz	
Oct	1	**EXAM 1**	
WEEK 4			
Oct	4	Lewis structures	Ch10.1
	6	VSEPR	Ch10.2
	7	Discussion, quiz	
	8	Shape & polarity	Ch10.3
WEEK 5			
Oct	11	Valence bond theory	Ch11.1-11.2
	13	Molecular orbital theory	Ch11.3
	14	Discussion, quiz	
	15	Water as a solvent	Ch4.1-4.2
WEEK 6			
Oct	18	Precipitation reactions	Ch4.3-4.4
	20	**EXAM 2**	
	21	Discussion, quiz	
	22	Oxidation-reduction reactions	Ch4.5-4.6
WEEK 7			
Oct	25	Acids and bases in water	Ch18.1-18.2
	27	Bronsted-Lowry	Ch18.3-18.4
	28	Discussion, quiz	
	29	Weak bases/weak acids	Ch18.5

WEEK 8			
Nov	1	Redox Reactions	Ch21.1
	3	Voltaic cells	Ch21.2-21.3
	4	Discussion, quiz	
	5	Forms of energy	Ch6.1-6.2
WEEK 9			
Nov	8	Calorimetry	Ch6.3-6.4
	10	Hess's Law	Ch6.5-6.6
	11	Discussion, quiz	
	12	**EXAM 3**	
WEEK 10			
Nov	15	Entropy	Ch20.1-20.2
	17	Calculating entropy change	Ch20.2-20.3
	18	Discussion, quiz	
	19	**FINAL EXAM** (comprehensive), 8-9:50 am, Olin 205	